# Types of Reactions Predicting Product Practice

Please note that the following reactions are **NOT necessarily** balanced! We are just focusing on types of reactions and predicting products. You can balance them later!

Things to Remember – crossing over to make neutral compounds, diatomics, roman numerals, rewriting water as H(OH), write good formulas and then balance at the end to fix your numbers!

#### Page Set Up

Q#	Type	Reaction
1		
2		
3		
Etc		



## BaCl<sub>2</sub> + H<sub>2</sub>O ->

#### NaBr >

$$(NH_4)(NO_3) + Ba(OH)_2 \rightarrow$$



### CaCO<sub>3</sub> + HCl <del>-</del>



#### K+HCI+



$$Ca(OH)_2 \rightarrow$$
 calcium oxide + water



## Na<sub>2</sub>S + H<sub>2</sub>O →



$$C_6H_8 + O_2 \rightarrow$$

### Al + CuSO<sub>4</sub> →



$$K + N_2 \rightarrow$$



#### Cu + AlCl<sub>3</sub> ÷

assume copper (II) in your product



#### CuSO<sub>4</sub> + NaOH →

assume copper (I) is in your product







### Fe + Cl<sub>2</sub> →

assume iron (IV) is in your product



### CaBr<sub>2</sub> + H<sub>2</sub>O ->



### Al+HNO<sub>3</sub> ÷

#### MgCO<sub>3</sub> + HCl →



$$C_2H_4 + O_2 \rightarrow$$

 $CO_2 + H_2O$ 

### Ca + CuCl<sub>3</sub> +



$$Ca(OH)_2 + AlBr_3 \rightarrow$$

$$ZnSO_4 + Ca(OH)_2 \rightarrow$$

#### Zn + AIN ->



### $MgCl_2 + H_2O \rightarrow$

### Zn + CuCl<sub>2</sub> →



$$C_6H_{10} + O_2 \rightarrow$$

### NaF + H<sub>2</sub>O ->



$$Na_2O_2 \rightarrow sodium oxide + oxygen$$



$$C_4H_8 + O_2 \rightarrow$$



#### $Cl_2 + NiI_2 \rightarrow$

assume nickle(I) is in your product

#### CuO >



#### $CaCO_3 \rightarrow$ calcium oxide + carbon dioxide

### AgNO<sub>3</sub> + NaCl →



$$C_2H_2 + O_2 \rightarrow$$

$$Ca + Cl_2 \rightarrow$$



## Br<sub>2</sub> + KF ->